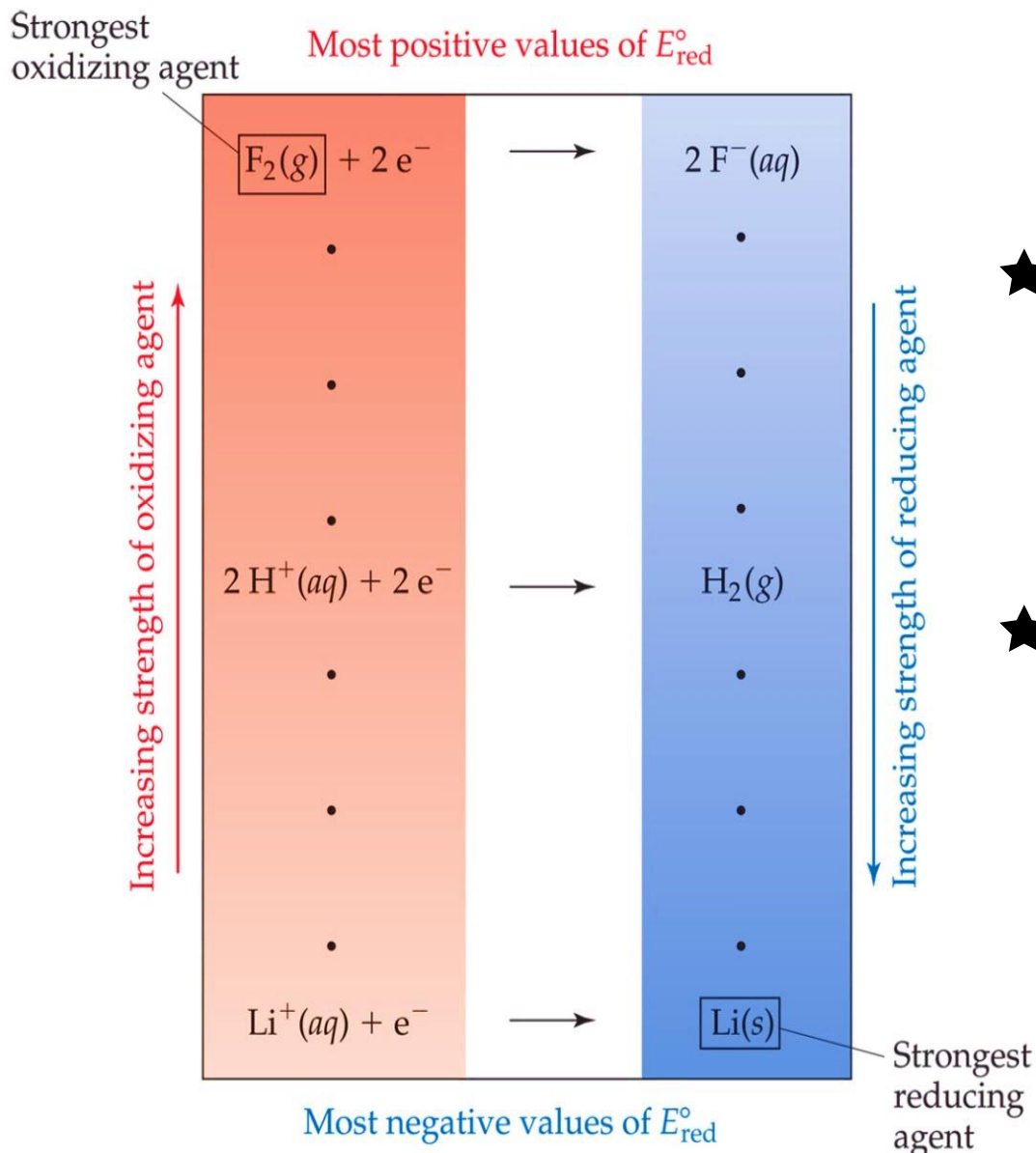


RELATIVE OXIDISING AND REDUCING AGENT STRENGTH



★ More positive E°

- More easily electron is gained
- More easily reduced
- Stronger oxidizing agent

★ More negative E°

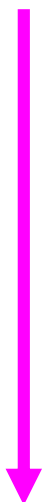
- More easily electron is lost
- More easily oxidised
- Stronger reducing agent

Oxidising agent

Reduction

E°_{red} (V)

Increasing strength of oxidising agent



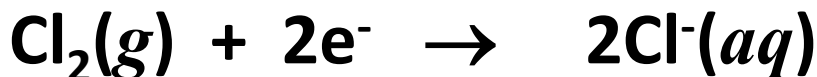
⋮

⋮



⋮

⋮



-0.25

⋮

0.00

⋮

+1.36

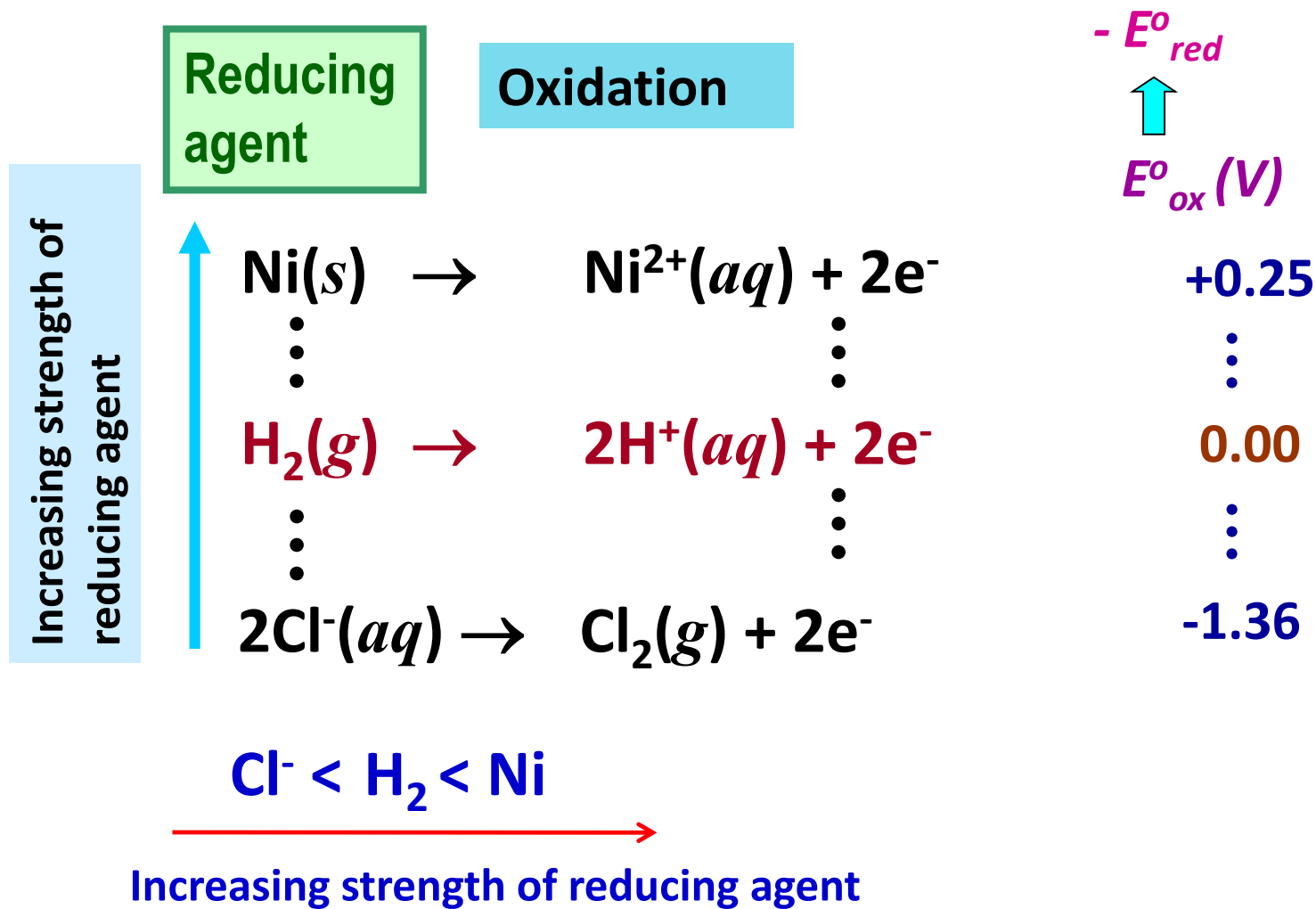
More +ve



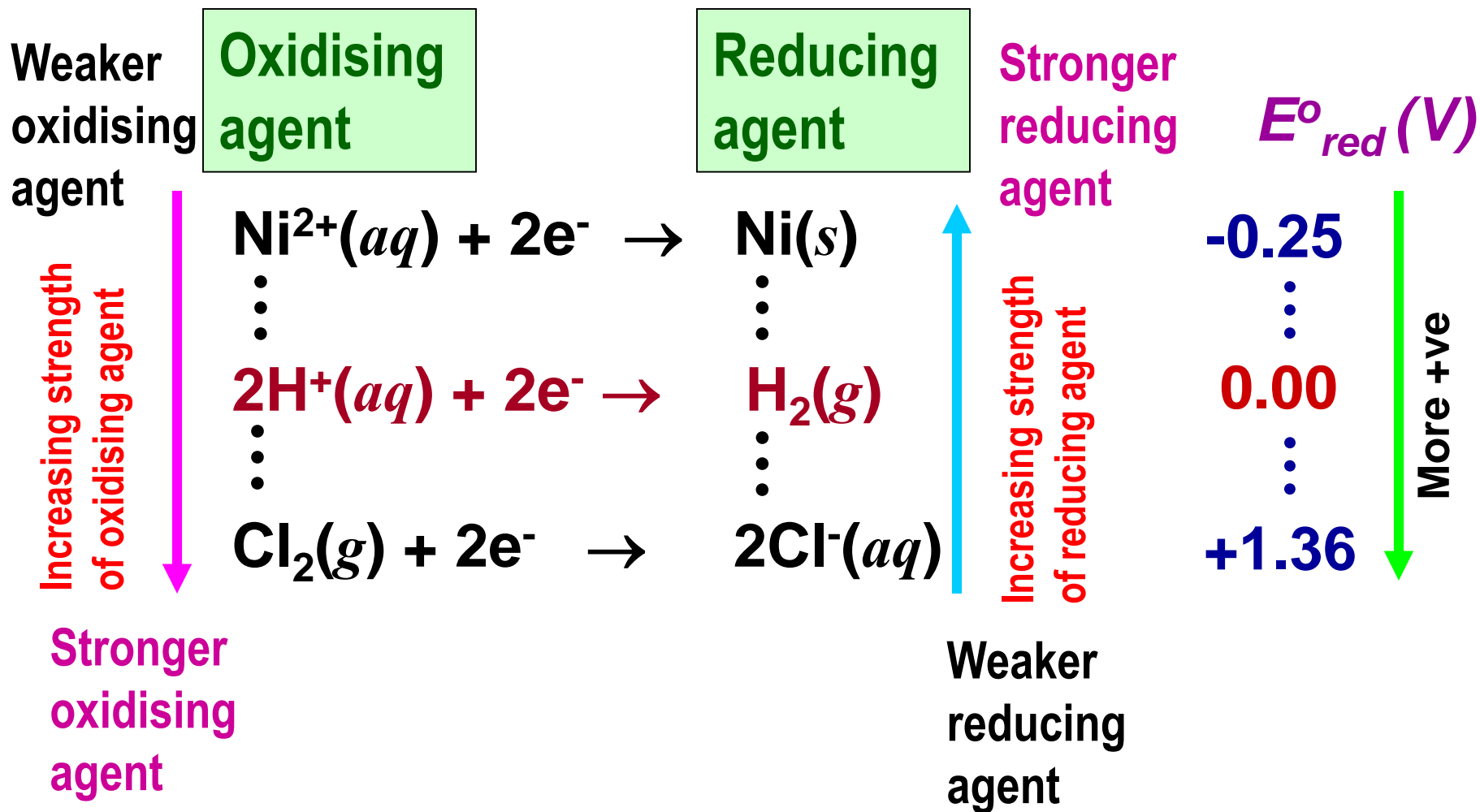
Increasing strength of oxidising agent

The more positive E°_{red} , the stronger the oxidising agent.

Note : For a reducing agent to supply electrons, the oxidation reaction must occur in the reverse direction.



The more negative E°_{red} , the stronger the reducing agent



- Based on the list of S.R.P, an **oxidising agent** (species on the **left**) that has a **larger value of E°_{red}** will react spontaneously with a **reducing agent** (species on the **right**) which has a **lower value of E°_{red}** .